Chapter 3 States of Matter

3.1 Solids, Liquids, and Gases (pg 68-74)

3.2 The Gas Laws (pg. 75-81)

3.3 Phase Changes (pg. 84- 91)

**Vocabulary**

|  |  |  |
| --- | --- | --- |
| **3.1 Vocab** | **3.2 Vocab** | **2.3 Vocab** |
| SolidLiquid Gas Kinetic Energy | Pressure Absolute Zero Charles’s LawBoyle’s Law | Phase ChangeEndothermic Heat of FusionExothermic VaporizationEvaporationVapor Pressure CondensationSublimationDeposition |

**Key Concepts**

(Questions you should be able to answer before exam.)

|  |  |  |
| --- | --- | --- |
| 3.1 Classifying Matter | 3.2 Physical Properties | 3.3 Chemical Properties |
| * How can shape and volume be used to classify materials?
* How can kinetic theory and forces of attraction be used to explain the behavior of gases, liquids, and solids?
 | * What causes gas pressure in a closed container?
* What factors affect gas pressure?
* How are the temperature, volume, and pressure of a gas related?
 | * What are six common phase changes?
* What happens to a substance’s temperature and system’s energy during a phase change?
* How does the arrangement of water molecules change during melting and freezing?
* How are evaporation and boiling different?
 |

**Extra Practice**

Section 3.1- Pg.74 Reviewing Concepts and Critical Thinking #1-8

Section 3.2- Pg. 81 Reviewing Concepts and critical thinking #1-8

Section 3.3- Pg. 91 Reviewing Concepts and Critical Thinking #1-8