Chapter 3 States of Matter

3.1 Solids, Liquids, and Gases (pg 68-74)

3.2 The Gas Laws (pg. 75-81)

3.3 Phase Changes (pg. 84- 91)

**Vocabulary**

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| **3.1 Vocab** | **3.2 Vocab** | **2.3 Vocab** |
| Solid  Liquid  Gas  Kinetic Energy | Pressure  Absolute Zero  Charles’s Law  Boyle’s Law | Phase Change  Endothermic  Heat of Fusion  Exothermic  Vaporization  Evaporation  Vapor Pressure  Condensation  Sublimation  Deposition |

**Key Concepts**

(Questions you should be able to answer before exam.)

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| 3.1 Classifying Matter | 3.2 Physical Properties | 3.3 Chemical Properties |
| * How can shape and volume be used to classify materials? * How can kinetic theory and forces of attraction be used to explain the behavior of gases, liquids, and solids? | * What causes gas pressure in a closed container? * What factors affect gas pressure? * How are the temperature, volume, and pressure of a gas related? | * What are six common phase changes? * What happens to a substance’s temperature and system’s energy during a phase change? * How does the arrangement of water molecules change during melting and freezing? * How are evaporation and boiling different? |

**Extra Practice**

Section 3.1- Pg.74 Reviewing Concepts and Critical Thinking #1-8

Section 3.2- Pg. 81 Reviewing Concepts and critical thinking #1-8

Section 3.3- Pg. 91 Reviewing Concepts and Critical Thinking #1-8